Double Reduction of 4,4-Bipyridine and Reductive Coupling of Pyridine by Two Thorium(III) Single-Electron Transfers

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Double Reduction of 4,4′-Bipyridine and Reductive Coupling of Pyridine by Two Thorium(III) Single-Electron Transfers


Abstract: The redox chemistry of uranium is burgeoning and uranium(III) complexes have been shown to promote many interesting synthetic transformations. However, their utility is limited by their reduction potentials, which are smaller than many non-traditional lanthanide(II) complexes. Thorium(III) has a greater redox potential so it should present unprecedented opportunities for actinide reactivity but as with uranium(II) and thorium(II) chemistry these have not yet been fully realized. Here we present reactivity studies of two equivalents of [Th(Cp′′)]2 [1, Cp′′ = {C12H14(SiMe3)12}–1,3)] with 4,4′-bipyridine or two equivalents of pyridine to give [[Th(Cp′′)]2(μ-NC(H2)2)] (2) and [[Th(Cp′′)]2(μ-(NC(H2)3)] (3), respectively, which contain doubly reduced substrates. As relatively large reduction potentials are required to effect these transformations we have shown that thorium(III) can promote reactions that uranium(III) cannot, opening up promising new reductive chemistry for the actinides.

Deepening our understanding of the actinides is crucial for the future development of bulk processes associated with nuclear fuel cycles.[1] For example, improved knowledge of reduction potentials developed during the Manhattan project led to the patenting of the PUREX (Plutonium Uranium Redox Extraction) process for selective plutonium extraction.[2] Whilst the reduction potentials of actinides in acidified aqueous solution are well established,[3] these differ markedly for molecular systems and this area is poorly developed for molecular thorium.[4] Thus, the reduction potentials of actinides in acidified aqueous solution are well established[3] and this area is poorly developed for molecular thorium.[4]

The reduction of these substrates has not previously been mediated by any U(II) complex.[5] However, Ln

3

Ln = lanthanide) chemistry is now blossoming[11] and the reductive coupling of N-heterocycles by Ln

4,4

Ln = Sm, Tb, Yb) complexes has been studied in depth previously.[12] It is noteworthy that Berthet, Eptikhine and co-workers have shown that U2 can reduce 2,4,6-tris(2-pyridyl)-1,3,5-triazine[13] and [U(Cp′)]2 (Cp′ = {C12H14(SiMe3)}) can reduce pyrazine.[14] Germane to this study Th(IV) complexes containing di-reduced 2,2′-bipy have been reported by Walter and Zl[15,16] and Arnold,[17] and uranium complexes containing mono-reduced 2,2′-bipy have been reported by Bart[19] and Cummins.[20]

Two equivalents of 1 doubly reduce 4,4′-bipy and reductively couple two equivalents of pyridine to yield [[Th(Cp′′)]2(μ-(NC(H2)2)] (2) and [[Th(Cp′′)]2(μ-(NC(H2)3)] (3), respectively (Scheme 1). Although selected Ln (Ln = Sm, Tb)[12c–e] and Sc[11] arsine[21] complexes can reductively couple pyridine, Eptikhine and co-workers have shown by single crystal XRD studies that the comparable uranium(III) complexes [U(Cp′)]2 and [U(Cp′)]2 (Cp′ = {C12H14(SiMe3)}) do not reduce pyridine[14,22] or 4,4′-bipy.[23] The formulations of 2 and 3 were confirmed by single crystal XRD (Figures 1–2), 1H, 13C{1H} and 29Si{1H} NMR spectroscopy and elemental analysis (see Supporting Information). The only structurally authenticated thorium complexes containing 4,4′-bipy are [Th(C2H4){2,4,6-bpy}][24] [Th(OTer5){2,4,6-bpy}][24] and [Th(OTer5){2,4,6-bpy}][24] and [Th(OTer5){2,4,6-bpy}][24] which have neutral 4,4′-bipy units. It is noteworthy that Th(IV) complexes containing di-reduced 2,2′-bipy previously reported in the literature[15–18] have to date exclusively been prepared by reduction of Th(IV) precursors with KC8 (E1/2 in DMF/0.1 M N[Bu4N]+ vs. Ag/AgCl: 2,2′-bipy = –2.19 and –2.76 V)[10] Complex 1 does not react with polyaromatics with smaller reduction potentials such as anthracene (E1/2 = –1.98 V) or naphthalene (E1/2 = –2.60 V)[26] indicating that coordination is necessary for electron transfer.

Scheme 1. Synthesis of 2 and 3 from 1: i) 0.5 eq 4,4′-bipy, toluene, 18 h; ii) pyridine, toluene, 5 d.

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Supporting information for this article is available on the WWW.
The UV-Visible spectra of 2 and 3 (0.1 mM in toluene) were collected between 200-1400 nm (see Supporting Information). Complex 2 exhibits a strong broad absorbance centered at 444 nm (ε = 19,400 M⁻¹cm⁻¹), assigned as a n→n* transition in the conjugated π-system, which is absent in 3. A sample of 3 (0.5 mM in toluene) was left to stand for four days at ambient temperature to develop an intense blue colour. The UV-vis spectrum of this sample was collected and the absorptions are comparable with an authentic sample of 1. [6]

Powder X-band and Q-band EPR spectra of 3 were collected at 298 K (see Supporting Information). These spectra are consistent with a radical anion ligand complex of Th⁺ with very small g-anisotropy (g∥ = 2.004, g⊥ = 1.990), which we tentatively attribute to the presence of "Th(Cp)₃(py)⁺". No additional features appeared at 5 K. In agreement with the UV-Visible spectra, when a toluene solution of 3 is left to stand for >30 mins the X-band EPR spectra were consistent with the formation of a Th⁺ complex, with g-values (g∥ = 1.975 g⊥ = 1.877) that correspond with data previously reported for EPR spectra of 1 in methylcyclohexane collected from 10–100 K (g∥ = 1.9725, g⊥ = 1.879). [6]

As an equilibrium mixture of compounds forms almost immediately (<10 mins) when 3 is dissolved in [D₆]toluene or [D₆]benzene we could not extract kinetic parameters for the reverse reaction by ¹H NMR spectroscopy. However, an equilibrium constant, Kₑ, at 298 K was determined after the solution was left for five days at room temperature to equilibrate by comparing the ¹H NMR integrals of pyridine with those attributed to arise from 3 (Kₑ = 0.5 mol dm⁻³). The value of this constant did not change when the ¹H NMR spectrum was collected at 303 K or 313 K. In an attempt to show that "Th(Cp)₃(py)⁺" is an intermediate in the formation of 3 a [D₆]benzene solution of 1 containing 2 eq. pyridine was treated with excess 1,4-cyclohexadiene. No benzene or H₂ was observed in the ¹H NMR spectra of the reaction mixture and after several days crystals of 3 formed, so an organic radical intermediate could not be unequivocally proven (see Supporting Information). Finally, no products could be identified from the treatment of 1 with 2 eq. 4-tert-butylpyridine in [D₆]benzene using ¹H NMR spectroscopy. This reaction mixture retained the intense blue colour associated with solutions of 1, even upon heating to 323 K for 16 hours (see Supporting Information).

The reversible C–C bond formation in 3 is analogous to the reactions of the Fe⁺ complexes [Fe(ArNCMe)₂]⁺(C₆H₆) (Ar = C₆H₅Pr₂-2,6) or [Fe(ArNCMe)₂]⁺(μ-N₂) with pyridine. Both reactions gave an Fe⁺ product containing 4,4'-dihydridobis(4,4'-pyridine)-1,1'-diyl, [[Fe₂(μ-N₂)(μ-ArNCMe)₄]]⁺, which exhibits a monomer-dimer equilibrium in solution. [27] Pyridine reduction by highly reducing Ln⁺ systems is well-established [21] but reversible C–C bond formation of an f-element complex containing a reduced unsubstituted pyridine ligand has not previously been observed to the best of our knowledge, although this has been seen for other N-heterocycles. [28] Unsurprisingly, 2 does not react with pyridine to give 3 as the reduction potential of 4,4'-bipy is less negative than that of pyridine and 1. [10] In contrast, we found that a [D₆]benzene solution of 3 reacts sluggishly with 4,4'-bipy to give 2 and pyridine. The slow rate of

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The ¹H NMR spectrum of 2 is simple and diagnostic of a symmetrical molecule, exhibiting a singlet at 0.45 ppm for the twelve SiMe₃ groups, two signals at 7.18 ppm and 6.98 ppm for the C₆H₇ protons and two broad signals at 6.80 ppm and 5.93 ppm for the two proton environments in the (μ-(NCMe)₃)₂ moiety. In contrast with 2, the ¹H NMR spectrum of 3 in [D₆]benzene is complex as it contains an equilibrium mixture of 3, pyridine and a postulated organic radical-bound Th⁺ intermediate complex, "[Th(Cp)₃(py)⁺]". Samples of 3 in aromatic solvents at room temperature turn pale blue within 30 minutes and intensely blue within 24 hours. This indicates the formation of a Th⁺ complex such as 1 as Th⁺ solutions tend to exhibit an intense blue colour. [8] Analysis of ¹H NMR spectra over several days showed that the relative amount of pyridine in these mixtures also tends to increase over time, so additional experiments were performed to analyze this equilibrium (see below). We found that dissolution of 3 in [D₆]benzene in the presence of a trace amount of pyridine slowed the reverse reaction sufficiently for the NMR spectra to be tentatively assigned. The signals due to 3 in the ¹H NMR spectrum of this mixture were assigned as follows: a singlet at 0.48 ppm for the twelve SiMe₃ groups, signals at 7.02 and 7.23 ppm for the C₆H₇ protons, and signals at 3.89, 4.95 and 6.91 ppm for the protons at the 4-, 3- and 2- positions, respectively of the 4,4'-dihydridobis-(4,4'-pyridine)-1,1'-diyl ligand. The ¹³C(¹H) NMR spectra of 2 and 3 are unremarkable, save for the inter-ring carbons which resonate at 149.56 and 44.13 ppm, respectively.
this reaction can be attributed to the equilibrium of 3 with 1 and pyridine, with competitive binding of pyridine and 4,4′-bipy to the single vacant coordination site of 1 (see Supporting Information).

A notable feature of the structure of 2 is the diagnostic inter-pyridyl C=C double bond length [1.376(10) Å], which is significantly shorter than those seen in 4,4′-bipy adducts such as [(U(Cp)₃)₂(4,4′-bipy)][1.474(17) Å][23] or [Th(C₅H₆)₂(4,4′-bipy)][1.478(4) Å][24] and is comparable to the distances in other direduced 4,4′-dihydroidbis(4,4′-pyridine)-1,1′-diyl complexes[25] and bis(trimethylsilyl)dihydro-4,4′-bipy [C=C: 1.381(3) Å][26a] The intra-pyridyl bond lengths are also diagnostic of direduction. The inter-ring C–C distance in 3 [1.562(12) Å] is typical of a single bond and this, together with other heterocyclic ring metrics, is comparable with similar complexes in the literature.[23c–e,21,27] The Th–N distances in 2 [2.359(4) Å] and 3 [2.350(4) Å] are shorter than those in Th₃N 4,4′-bipy adducts [range: 2.626(2)-2.707(2) Å][24,25] and are typical of Th³⁻–N amide bonds, e.g. [Th[N(SiMe₃)₂]₂(BH₄)] [2.32(2) Å].[26b]

The electronic structures of 2 and 3 were characterized at the density functional theory (DFT) level, employing the PBE0 exchange-correlation functional and a polarized split-valence basis set for structural optimizations (see Supporting Information for full details). Electronic properties were derived from single-point energy calculations using a polarized valence triple-ζ basis set. Structural parameters were in excellent agreement with experiment, with bond lengths typically deviating from experimental values by less than 0.02 Å (see Supporting Information Table S4), justifying the model chemistry. Further confidence in these models was provided by the bulk features of the UV/visible spectrum of 2 and the IR spectra of 2 and 3 being reproduced with reasonable agreement by calculated values (see Supporting Information Tables S5–7). Inspection of the resultant molecular orbitals reveals the presence of both σ- and π-type Th–N bonding interactions (Figure 3), although the thorium contributions are, as would be expected, small.

The Th–N bond orders, obtained via the quantum theory of atoms in molecules (QTAIM)/Mayer approaches, were calculated to be 0.558/0.767 and 0.573/0.798 for 2 and 3, respectively. NBO analysis failed to identify any Th–N bonding orbitals, presumably due to Th contributions falling below the 5% default threshold, but Th contributions to the orbitals shown in Figure 3 were estimated via QTAIM and Hirshfeld partitioning of the molecular spaces. Both methods predicted small (~2-4%) Th contributions, and 3 was further investigated using QTAIM (See Supporting Information Table S9). All metrics indicate a predominantly ionic Th–N interaction with very similar covalent character. QTAIM metrics suggest a slightly more covalent interaction in 3. This is commensurate with the higher bond order and shorter Th–N bond: combined these measures indicate a stronger Th–N bond in 3. Th–N bond rigidities were calculated to be 0.29 and 0.27 for 2 and 3, respectively. When compared to those of benzene (0.23) and ethylene (0.45) these values indicate a degree of double bond character in the Th–N bonds of 2 and 3.

To conclude, we have demonstrated that two equivalents of the Th³⁻ complex 1 promote the double reduction of 4,4′-bipyridine and the reductive coupling of pyridine. This work shows that Th³⁻ complexes can exhibit reductive chemistry of the order of non-traditional Ln³⁻ systems, opening up new reductive chemistry for the actinides. The consequence is that the reductive small molecule activation chemistry of Th³⁻, thus far burgeoning only for U³⁻ in the actinide series, should yield contrasting and fascinating results in future.

**Experimental Section**

Full synthetic details, characterization data and computational data for 2-3 is available in the Supporting Information. Additional research data supporting this publication are available from The University of Manchester eScholar repository at DOI:10.15127/1.302738.

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**Keywords:** thorium • N ligand • subvalent compounds • reduction • electron transfer


**The Mighty Thorium:** The reductive chemistry of the thorium(III) complex, 
\([\text{Th}(\text{C}_2\text{H}_3\text{SiMe}_3)_2\cdot\text{H}_2\text{O}]_3\), is shown to be of the order of non-traditional lanthanide(II) complexes, opening up new reductive chemistry for the actinides.

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